

# Download Equilibrium Constant Determination Of Iodide

For the answer to the question above, The rate law for the slow step is:  $\text{rate} = k [C] [D]$  Since [C] is an intermediate, we need to substitute it using equilibrium step. We know equilibrium means the rate of forwarding reaction = rate of reverse reaction. Flinn Scientific is the #1 source for science supplies and equipment both in and outside the classroom. For more than 40 years, Flinn has been the "Safer Source for Science." Eastern Michigan University Education First. Department of Chemistry. 541 Mark Jefferson Science Complex Ypsilanti, Michigan 48197 The rate law or rate equation for a chemical reaction is an equation that links the reaction rate with the concentrations or pressures of the reactants and constant parameters (normally rate coefficients and partial reaction orders). For many reactions the rate is given by a power law such as  $= [A]^m [B]^n$  where [A] and [B] express the concentration of the species A and B (usually in moles per liter ...